

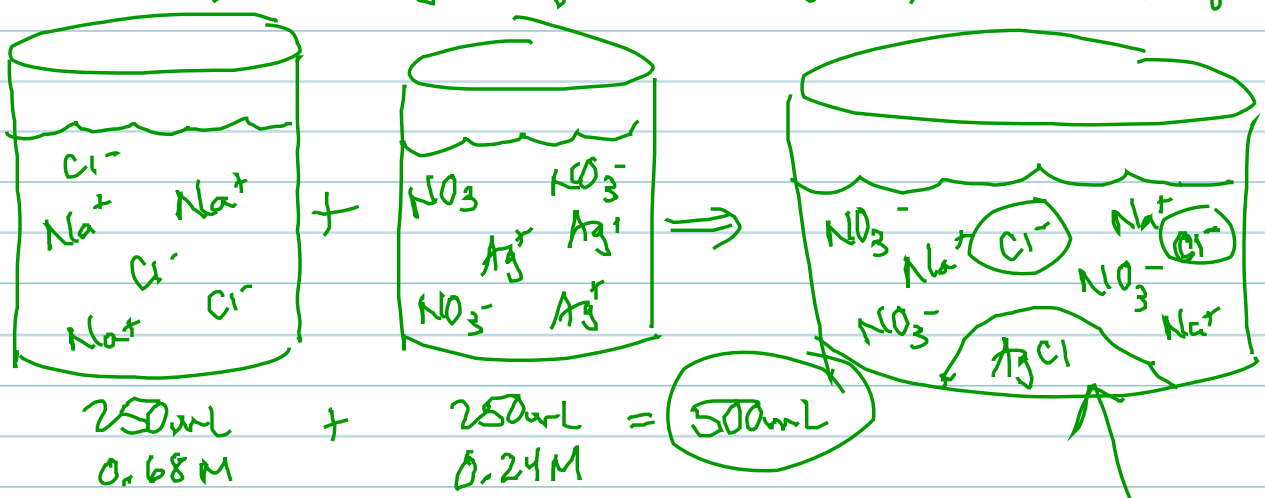
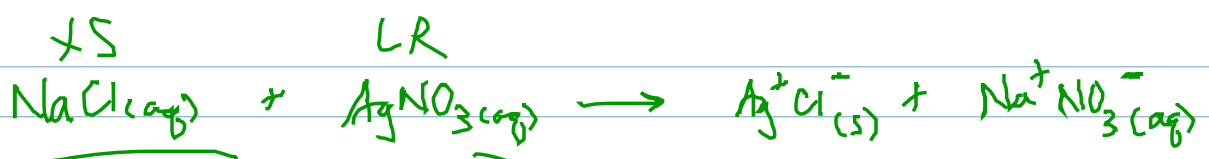
$$\text{NaCl: } (0.68\text{M})(0.25\text{L}) = 0.17\text{mol NaCl} \times \frac{1\text{mol AgCl}}{1\text{mol NaCl}} = 0.17\text{mol AgCl}$$

$$\text{AgNO}_3: (0.24\text{M})(0.25\text{L}) = 0.06\text{mol AgNO}_3 \times \frac{1\text{mol AgCl}}{1\text{mol AgNO}_3} = 0.06\text{mol AgCl}$$

$$0.06\text{mol AgCl} \times \frac{143.32\text{g}}{1\text{mol}} = 8.60\text{g AgCl}$$

$$\begin{array}{r} 0.17\text{mol AgCl} \\ - 0.06\text{mol AgCl} \\ \hline 0.11\text{mol AgCl} \end{array} \times \frac{1\text{mol NaCl}}{1\text{mol AgCl}} = \frac{0.11\text{mol NaCl}}{0.5\text{L}}$$

$\text{NaCl} \Rightarrow 0.22\text{M}$



$\text{NaCl: } (0.68\text{M})(0.25\text{L}) = 0.17 \text{ mol NaCl} \times \frac{1 \text{ mol AgCl}}{1 \text{ mol NaCl}} = 0.17 \text{ mol AgCl}$

$\text{AgNO}_3: (0.24\text{M})(0.25\text{L}) = 0.060 \text{ mol AgNO}_3 \times \frac{1 \text{ mol AgCl}}{1 \text{ mol AgNO}_3} = 0.06 \text{ mol AgCl}$

$0.06 \text{ mol AgCl} \times \frac{143.32\text{g}}{1 \text{ mol}} = 8.60\text{g AgCl}$

$\frac{0.17 \text{ mol AgCl} - 0.06 \text{ mol AgCl}}{0.11 \text{ mol AgCl}} \times \frac{1 \text{ mol NaCl}}{1 \text{ mol AgCl}} = \frac{0.11 \text{ mol NaCl}}{0.5\text{L}}$

$\text{NaCl} \Rightarrow 0.22\text{M}$

